

Nanoscale Engineering of Hybrid Magnetite-Carbon Nanofibre Materials for Magnetic Resonance Imaging Contrast Agents

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Abstract

Magnetic nanomaterials show significant promise as contrast agents for magnetic resonance imaging (MRI). We have developed a new highly efficient one-step procedure for the synthesis of magnetically-functionalised hollow carbon nanofibres, where (i) the carbon nanofibres act as both a template and a support for the nucleation and growth of magnetite nanoparticles and (ii) the structural (size, dispersity and morphology) and functional (magnetisation and coercivity) properties of the magnetic nanoparticles formed on nanofibres are strictly controlled by the mass ratio of the magnetite precursor to the nanofibres and the solvent employed during synthesis. We have shown that our magnetite-nanofibre materials are effectively solubilised in water resulting in a stable suspension that has been employed as a “negative” MRI contrast agent with an excellent transverse relaxivity (r_2) of (268 ± 13) mM/s, surpassing current commercial materials and state-of-the-art magnetic nanoscale platforms in performance for MRI contrast at high magnetic fields. The

preparation and evaluation of this unique hybrid nanomaterial represents a critical step towards the realisation of a highly efficient “smart” MRI theranostic agent – a material that allows for the combined diagnosis (with MRI), treatment (with magnetic targeting) and follow-up of a disease (with MRI) – currently in high demand for various clinical applications, including personalised nanomedicine.

Keywords: nanoparticle, nanofibre, carbon nanotube, magnetism, magnetic resonance imaging.

1. Introduction

Hollow one-dimensional carbon nanostructures, such as carbon nanotubes (CNTs) and graphitised nanofibres (GNFs), are chemically stable and mechanically robust and have been demonstrated as effective containers for functional molecules and nanoparticles.¹⁻⁵ Whilst much of the current literature has focussed on the applications of narrow CNT as containers, comparatively little attention has been paid to GNFs. These unique hybrid carbon nanostructures comprise an atomically-smooth multi-walled nanotube exterior and a cup-stacked interior which gives rise to internal graphitic protrusions, approximately 3.5 nm in height and spaced at 8-15 nm intervals along the nanofibre growth axis, where localised van der Waals forces are maximised and molecules and nanoparticles become spontaneously immobilised (**Figure 1**).⁶⁻⁸

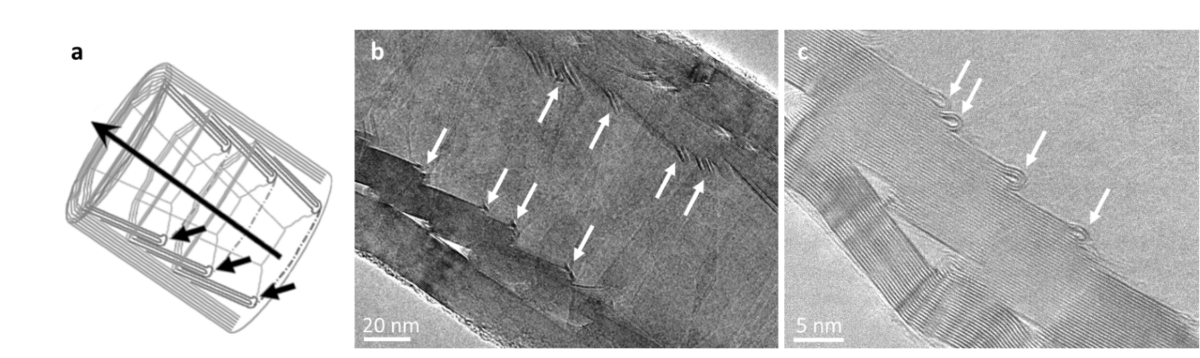


Figure 1. (a) Schematic representation and (b) and (c) transmission electron microscope (TEM) micrographs highlighting the unique structure of graphitised nanofibres. The step-edge arrangement,

denoted by the small black arrows in (a) and the white arrows in (b and c), provides an area for both molecules and nanoparticles species to be selectively deposited.

Unlike CNTs, which require opening and have an internal diameter of a few nanometres, graphitised nanofibres are always open and have comparatively large accessible internal volumes, often more than 50 nm in diameter, allowing for the rapid, uninhibited diffusion of molecules and nanoparticles into the nanofibres. These distinguishing properties of GNFs make them superior to CNT and thus ideal vessels for magnetic resonance imaging (MRI) contrast materials, such as magnetite nanoparticles ($\text{Fe}_3\text{O}_4\text{NP}$). Magnetite composites have shown significant promise for MRI methods due to their excellent magnetic characteristics, such as high magnetic saturation;⁹⁻¹¹ however, the magnetic properties of Fe_3O_4 are highly dependent on the size, shape, crystallinity and purity of magnetite nanoparticles and their distribution on a given substrate.^{12,13} The current challenge for magnetite-based MRI contrast agents is to engineer Fe_3O_4 nanostructures to maximise the rate of transverse relaxation (R_2) whilst maintaining a low concentration of magnetite in the system.

In this comparative study, we appraise the structural and magnetic properties of a series of magnetite nanoparticle/graphitised nanofibre composites ($\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}$) prepared by simple one-step procedures involving readily accessible reagents. Despite the fact that the formation of $\text{Fe}_3\text{O}_4\text{NP}$ on CNTs has been extensively studied previously,¹⁴⁻²⁷ surprisingly, little investigation has been performed on the formation of magnetite on and in GNFs to date.²⁸⁻³¹ Herein, we demonstrate that GNFs serve as effective templates for the *in-situ* formation of Fe_3O_4 nanoparticles. We show that reaction conditions, such as the ratio of GNF to iron precursor and the solvent system employed, have a significant effect on the size and morphology of Fe_3O_4 in the afforded GNF hybrids. Moreover, we observe that materials with larger, sparsely distributed, magnetite nanoparticles on GNF showed the best magnetic characteristics for MRI and thus the greatest potential for dual functionality theranostics.

2. Experimental

2.1 General Experimental

Graphitised nanofibres (PR19 Pyrograf Products, produced by chemical vapour deposition, mean length ~5 microns, mean inner and outer diameters are ~50 and ~120 nm respectively) were purchased from Applied Science, USA. All other reagents and solvents were purchased from Sigma-Aldrich, UK and used without further purification. Water was purified (> 18.0 MΩ cm) using a Barnstead NANOpure II system. High resolution transmission electron microscopy (HRTEM) was performed using a JEOL 2100F transmission electron microscope (field emission electron gun source, information limit 0.19 nm). Statistical analysis of nanoparticle diameter (d_{NP}) was conducted using GatanDigitalMicrograph software ($N > 200$). Energy dispersive X-ray (EDX) analysis was performed using an Oxford Instruments INCA 560 X-ray microanalysis system. TEM/EDX specimens were prepared by casting several drops of methanolic solutions onto copper-grid- or nickel-grid-mounted “holey” carbon films and drying under a stream of nitrogen. Thermogravimetric analysis was determined using a TA Instrument equipped with a SDT Q600 Analyser under flowing air at a rate of 10 °C min⁻¹ up to 800 °C. Raman spectroscopy was conducted using a Horiba–Jobin–YvonLabRAM HR spectrometer, with a laser wavelength of 532 nm operating at a power of *ca.* 4 mW and a 600 lines mm⁻¹ grating. The detector was a Synapse CCD detector. Spectra were collected by averaging 8 acquisitions of 60 s duration. The Raman shift was calibrated using the Raleigh peak and the 520.7 cm⁻¹ silicon line from an Si(100) reference sample. Measurement of the magnetic properties was performed using PPMS-9 (Physical Property Measurement System, EverCool II, Quantum Design) in magnetic fields up to 9 T. XRD measurements were carried out using a Rigaku Ultima IV X-ray diffractometer in the Bragg-Brentano recording geometry with Co_{Kα} radiation with a graphite monochromator on the diffracted beam. A Fisherbrand Accuspin centrifuge was used at 6080xg during purification of the composite materials.

2.2 Preparation of Fe₃O₄NP/GNF

The preparation of Fe₃O₄NP/GNF was achieved through modification of the procedure outlined by Chen *et al.*³² GNFs (25 mg) were heated at 400 °C for 1 h to remove any residual moisture from the internal channel. To this was added a solution of iron(III) acetylacetonate

(25 mg) in deionised water (6 mL) and the suspension sonicated at room temperature for 2 h. The solvent was removed *in vacuo* and the resulting solid heated from room temperature to 140 °C at a rate of 2.5 °C min⁻¹ and then held at 140 °C for 8 h in air. The solid was subsequently further heated to 350 °C at a rate of 2 °C min⁻¹ under a continuous flow of argon and held at 350 °C for an additional 3 h. Other materials reported in this work with differing ratios of Fe(acac)₃ to GNF and solvent (**Table 1**) were prepared in a similar fashion.

2.3 Biological applications

2.3.1 Dispersion of Fe₃O₄NP/GNF for biological applications

To Fe₃O₄NP/GNF-1:2w (5 mg) was added an aqueous solution of Pluronic F-127 (10 mL, 2 % v/v) and the mixture sonicated for 2 h at room temperature. After sonication, the dispersion was centrifuged (6,080xg, 20 min). Following centrifugation, the supernatant and precipitate were separated by decantation. The supernatant was discarded and the precipitate re-dispersed in the same volume of deionised water and centrifuged again. This procedure was repeated three times to yield a stable dispersion of Fe₃O₄/GNF:Pluronic in deionised water. The percentage of Pluronic F-127 in the final Fe₃O₄/GNF composite was determined by thermogravimetric analysis of the solid prior to the final re-dispersion in water and found to be 3.6 %.

Five 10 mL vials of Fe₃O₄NP/GNF-1:2w:Pluronic F-127 were prepared by dilution of the stock solution to five different concentrations for the use in the MRI experiments.

2.3.2 MRI experiments

MRI characterisation was performed using a 9.4 T Avance III micro-imaging system (Bruker, Germany). A 30 mm micro-imaging ¹H coil was used in all experiments. Standard MSME (Multi-Slice Multi-Echo) imaging protocol was used for simultaneous spatial mapping of either longitudinal (T₁) or transverse (T₂) time constants of water protons from five different aqueous dispersions of Fe₃O₄NP/GNF-1:2w:Pluronic F-127 containing 0.0024561, 0.0064311, 0.0128622, 0.0184211 and 0.0385866 mM of Fe ion. T₁ time constants were measured using saturation recovery method with 15 pre-pulses at time to echo (TE) of 9.7 ms. At least five

recycle delays were used to ensure full characterization of the mono-exponential longitudinal recovery. T_2 time constants were measured using TEs of 15, 50 and 100 ms and a recycle delay of 3000 ms. A 0.1 ms bipolar pulse was used for excitation and 1 ms Gaussian pulse was used for slice selective re-focusing. Data from central 5 mm axial slices were collected into 128x128 matrices with isotropic field of view (FOV) of 30 mm², resulting in raw spatial resolution of 0.234 mm². Data were zero filled to 256x256 before image reconstruction. Images were reconstructed using Prospa 3.1 (Magritek, New Zealand). T_1 and T_2 data analysis was performed using IgorPro 6.1 (Wavemetrics, USA). T_1 and T_2 time constants were determined in each phantom using pixel by pixel linear fitting after transforming the data into semi-logarithmic scale as mono-exponential proton relaxation behaviour was assumed. Subsequent determination of the averaged T_1 and T_2 values and their variances was performed using Gaussian multi-peak build-in subroutine.

3. Results and discussion

The thermal decomposition of iron(III) acetylacetonate ($\text{Fe}(\text{acac})_3$) was chosen as the ideal method for the preparation of Fe_3O_4 nanoparticles owing to its synthetic ease and use of cheap, widely accessible reagents. As $\text{Fe}(\text{acac})_3$ had to be both deposited on and encapsulated in GNFs to provide the precursor for the nanoparticles, it was, therefore, essential to control the physical properties of the solvent to ensure optimum deposition. We, thus, explored the use of both water and a mixture of water and ethanol for the effective solvation of $\text{Fe}(\text{acac})_3$. In a typical experiment, the metal complex was dissolved in water or a mixture of water/ethanol (3/1 v/v), GNFs were then immersed in the precursor solutions and mixed under ultrasonic agitation to ensure efficient mixing, followed by evaporation of the solvent to enhance the extent of adsorption of the precursor on the surfaces of GNFs. The GNFs and metal complex were then heated in air and finally in an inert atmosphere to promote the efficient decomposition of $\text{Fe}(\text{acac})_3$ to Fe_3O_4 .

Confirmation of the chemical composition and structure of the resultant nanoparticles in the hybrid structures was afforded by energy dispersive X-ray (EDX) analysis (**Figure 2a**), X-ray diffraction (XRD) (**Figure 2b**), Raman spectroscopy (**Figure 2c**) and high resolution transmission electron microscopy (HRTEM) (**Figure 2d**).

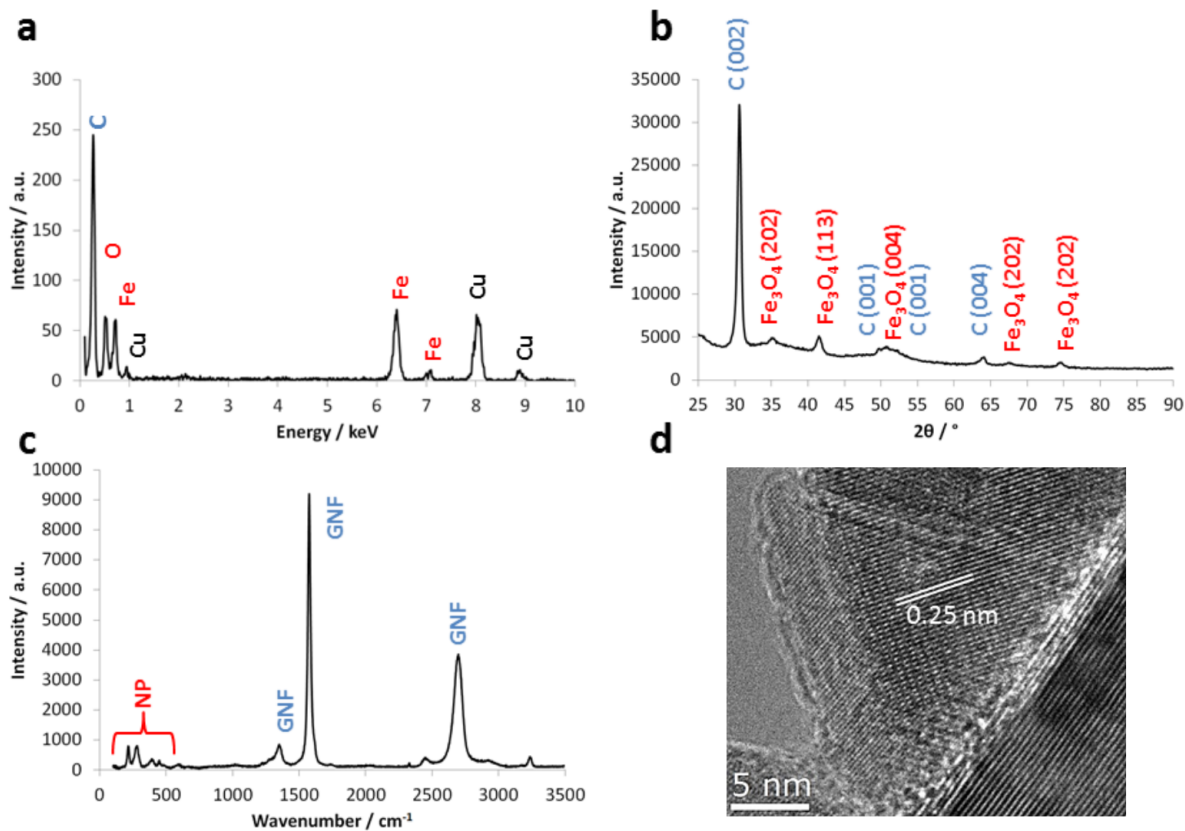


Figure 2. (a) A representative EDX spectrum of the Fe₃O₄NP/GNF hybrids. The peaks at 0.71, 6.40 and 7.06 keV correspond to L α , K α and K β fluorescence peaks of Fe respectively. The K α X-rays of O are observed at 0.53 keV. The atomic ratio of Fe:O of 3.0:4.2 matches well with that expected for Fe₃O₄, with the slight excess of oxygen attributed to a small degree of surface functionality of the GNF. The large peak at 0.28 keV is due to the K α X-rays of C in GNF. The large fluorescence peaks near 8.00 and 8.84 keV are the K α and K β X-ray fluorescence signatures of Cu respectively (an artefact induced from either the TEM grid mesh or the instrument column assembly). **(b)** A representative XRD pattern for Fe₃O₄NP/GNF hybrids showing the indexed crystal planes of both Fe₃O₄ and GNF. **(c)** A representative Raman spectrum of the Fe₃O₄-GNF composite showing bands associated with iron oxide (the peaks at 209, 273, 386 and 577 cm⁻¹ corresponding to one A_{1g} and three E_g modes in haematite, α -Fe₂O₃, a known by-product of the heated-induced transformation of Fe₃O₄ under the laser irradiation required for Raman spectroscopy analysis)³³ and GNF (1341, 1562 and 2684 cm⁻¹ corresponding to D, G and 2D modes in graphitic carbon structures respectively).³⁴ **(d)** A representative TEM micrograph of magnetite nanoparticles on the external GNF surface. An interplanar spacing of 0.25 nm can be observed that is consistent with the (311) Fe₃O₄ crystal plane.

Comparative TEM analysis revealed magnetite nanoparticles on GNF with average sizes of 11 ± 4 and 14 ± 3 nm for the materials prepared with the same mass ratio of $\text{Fe}(\text{acac})_3$:GNF in water and water/ethanol respectively (**Figure 3**).

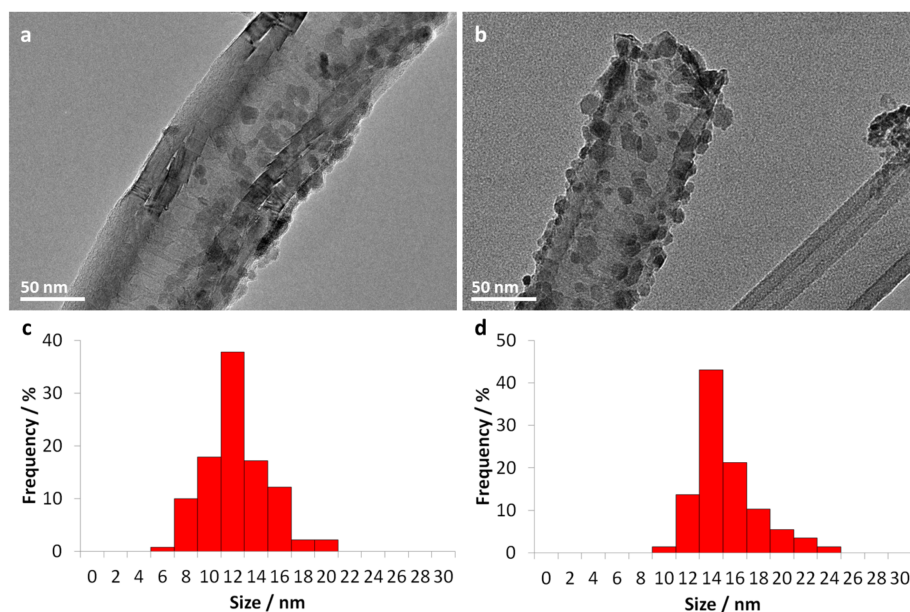


Figure 3. HR-TEM micrographs of $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}$ prepared using the same ratio of iron precursor to GNF ($\text{Fe}(\text{acac})_3$:GNF, 1:1) but using different solvent systems: **(a)** sample $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}$ -1:1w, where $d_{\text{NP}} = 11 \pm 4$ nm; **(b)** sample $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}$ -1:1ew where $d_{\text{NP}} = 14 \pm 3$ nm. The corresponding particle size distributions are displayed in **(c)** and **(d)** respectively.

Analysis of multiple micrographs revealed that a greater proportion of the Fe_3O_4 nanoparticles formed by the water/ethanol method appeared to reside on the external surface of the GNFs in comparison to the water method, where the number of nanoparticles confined within the inner channel was higher. Since the surfaces of GNFs are hydrophobic, water/ethanol wets the GNFs more effectively than pure water, allowing for a more even distribution of magnetite. The poor solubility of $\text{Fe}(\text{acac})_3$ in water, in conjunction with the hydrophobic properties of the GNFs, permitted a more efficient insertion and deposition of the precursor molecules into the internal channel of the GNFs.

More importantly, magnetic measurements in a 3 T field (**Figure 5**) demonstrated that the $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}$ hybrid material prepared using the same mass ratio of iron precursor to GNF

(Fe(acac)₃:GNF, 1:1) if made in water possesses a significantly lower magnetisation ($M_s = 51$ and 94 emu/g respectively) and a comparable coercivity ($H_c = 64$ and 31 Oe respectively) relative to Fe₃O₄NP/GNF prepared using water/ethanol (**Table 1**). The observed differences in magnetic behaviour can be readily explained by the distinct structural properties of the hybrid materials. In Fe₃O₄NP/GNF-1:1w, the mean diameter of nanoparticles is 11 ± 4 nm (above the critical diameter for magnetite to exist in a superparamagnetic state of $d_{sp} \lesssim 9$ nm)³⁵ and the density of the nanoparticles on GNF is low (only 140 per 1 μ m of GNF length). Consequently, there are only weak magnetostatic interactions between neighbouring nanoparticles, with approximately 50% of the nanoparticles existing in a superparamagnetic state even in the 3 T field as evident from the measured magnetisation which, therefore, do not contribute to the total magnetisation of the hybrid material.

As a consequence, water was used to prepare a series of samples to explore the effects of the ratio of the magnetite precursor to GNFs. Using different starting ratios of Fe(acac)₃ to GNFs allows for the control of not only the size and shape of Fe₃O₄ nanoparticle, but also potentially their distribution along the GNF and relative positions with respect to each other. More widely distributed and slightly larger nanoparticles are expected to have a larger surface of interaction with surrounding water molecules in MRI experiments, which should yield better contrast. We investigated three different mass ratios of Fe(acac)₃:GNF, 1:2, 1:1 and 2:1, showing that the average nanoparticle size varies from 8 to 14 nm across the series (**Figure 4**).

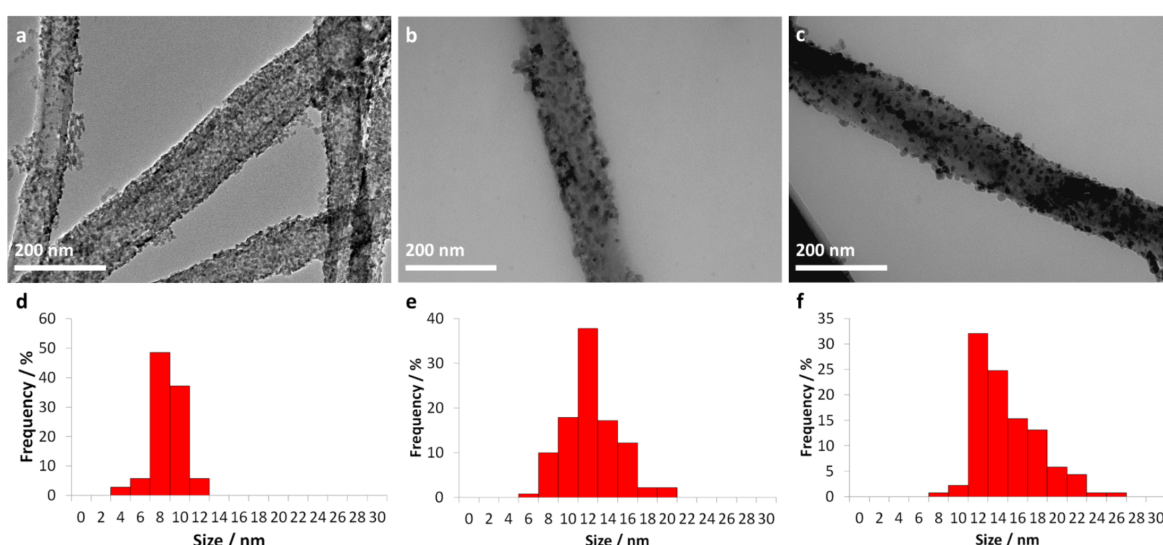


Figure 4. HR-TEM micrographs of Fe₃O₄NP/GNF prepared using three different mass ratios of Fe(acac)₃:GNF: **(a)** Fe₃O₄NP/GNF-2:1w, where $d_{NP} = 8 \pm 1$ nm; **(b)** Fe₃O₄NP/GNF-1:1w, where $d_{NP} = 11 \pm 4$ nm and **(c)** Fe₃O₄NP/GNF-1:2w, where $d_{NP} = 14 \pm 3$ nm, using water as the solvent. The corresponding particle size distributions are displayed in **(d)**, **(e)** and **(f)** respectively.

There is a clear relationship between the average nanoparticle size and the ratio of Fe(acac)₃:GNF. We observe that the higher the ratio of the precursor, i.e. a greater number of nucleation centres on GNFs, leads to a larger number of particles with a smaller average size and, therefore, a narrower particle size distribution (**Table 1**).

Table 1. Conditions used for the synthesis and the resultant properties of the Fe₃O₄NP/GNF hybrid materials.

Material	Fe(acac) ₃ : GNF (w:w)	Solvent	d_{NP} / nm	Density / NP μm^{-1}	Magnetic moment / emu g^{-1}	Coercivity, Oe
Fe ₃ O ₄ NP/GNF-1:2w	1:2	water	14 ± 3	310	99	173
Fe ₃ O ₄ NP/GNF-1:1w	1:1	water	11 ± 4	140	51	64
Fe ₃ O ₄ NP/GNF-2:1w	2:1	water	8 ± 1	>1000	83	82
Fe ₃ O ₄ NP/GNF-1:1ew	1:1	ethanol / water	14 ± 3	370	94	51

Most significantly, Fe₃O₄NP/GNF prepared in a 1:2 mass ratio has nanoparticles that are well dispersed (**Figure 4c**) due to the large surface of the GNF. Furthermore, evaluation of the magnetic properties of these materials yielded key insight into the structure-property relationships of nanoscale magnetic materials. For example, Fe₃O₄NP/GNF-2:1w possesses nanoparticles of a mean diameter which is comparable or below the critical diameter d_{sp} required for the transition of magnetite to the superparamagnetic state and can thus explain the high magnetisation of this material. Furthermore, a high density of nanoparticles on the surface of GNF in this material (>1000 per 1 μm of GNF length) provides strong magnetostatic interactions between them, which can be further enhanced by very effective dispersion

interactions between Fe₃O₄NP and GNF leading to an increase of the surface magnetic anisotropy parameter. All these phenomena can lead to an increase of the magnetic anisotropy constant and ‘freezing’ of the magnetic moments of many nanoparticles in this material (the estimated volume fraction of nanoparticles existing in the superparamagnetic state is below 10%). Overall, both the magnetization and the coercive force of Fe₃O₄NP/GNF-2:1w are enhanced as a result. Conversely, Fe₃O₄NP/GNF-1:2w possess relatively large nanoparticles, with a comparatively lower density of nanoparticles on the surface of GNF, leading to weaker magnetostatic interactions and thus more attractive magnetic properties (Figure 5).

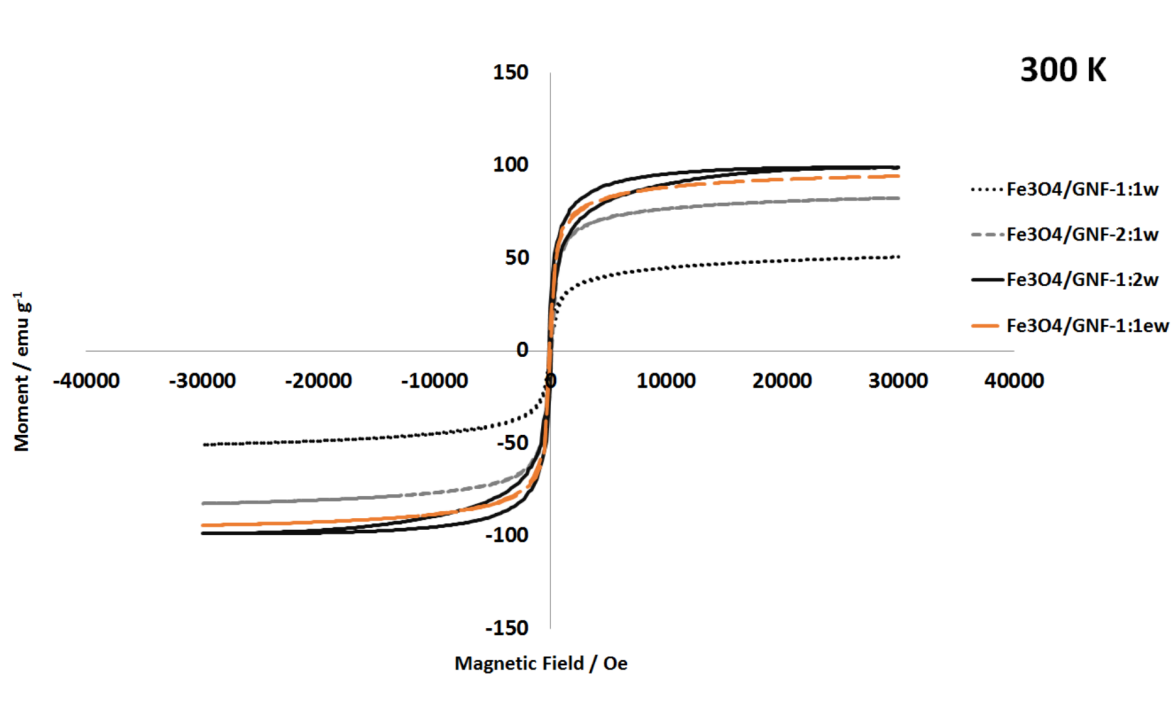


Figure 5. Magnetic properties of Fe₃O₄NP/GNF hybrids with varying ratios of Fe(acac)₃:GNF and solvent systems (w – water, ew – ethanol/water).

Since the hybrid Fe₃O₄NP/GNF material, prepared in water using a 1:2 ratio of Fe(acac)₃:GNF, possesses the most optimum structure and magnetic properties, it was tested as a contrast agent for MRI. However, in order to stabilise Fe₃O₄NP/GNF in water, a fundamental requirement for biological applications, Pluronic F-127 was added to the composite material, with the extent of coating confirmed with thermogravimetric analysis to be 11 % by mass.

The efficiency of MRI contrast agents depends on their longitudinal (r_1) and transverse (r_2) relaxivities, as well as on their ratio. The efficiency of a T_2 contrast agent is optimal when r_2/r_1 ratio is the highest. To determine the r_2/r_1 ratio in the presence of $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}$, the transverse relaxation (T_2) and longitudinal relaxation (T_1) of water protons in aqueous solutions of solubilised $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}$ were studied for a range of concentrations of the added contrast agent. T_1 and T_2 weighted images were obtained as described in the Experimental section and T_2 and T_1 relaxation maps derived from the resulting images (**Figure 6a and c**). The inset in Figure 6 lists the Fe cation concentrations in five aqueous samples of $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}$ that were used as an imaging phantom.

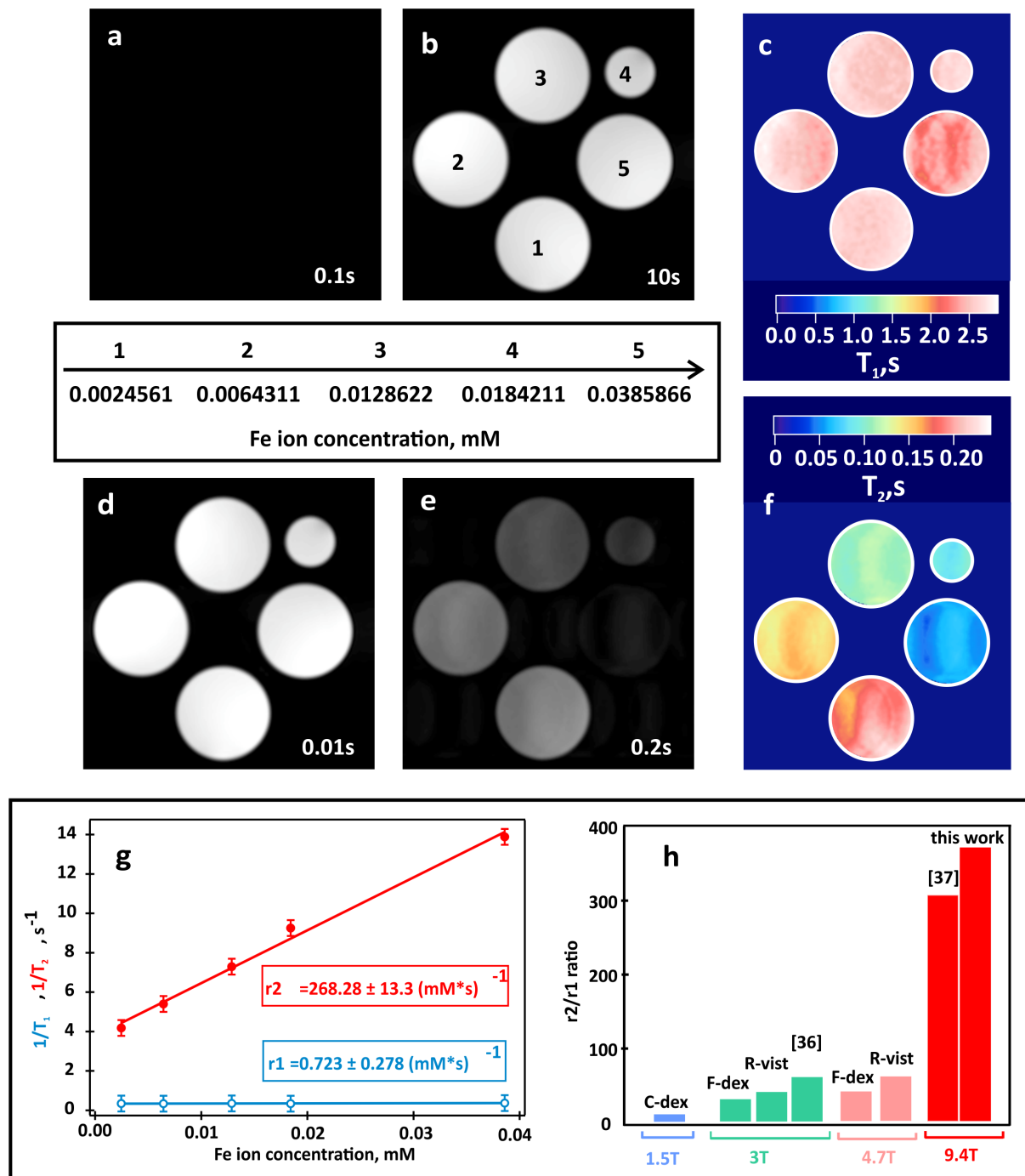


Figure 6. Longitudinal (r_1) and transverse (r_2) relaxivity of water in the presence of $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}-1:2\text{w}:\text{Pluronic F-127}$ at 9.4T: (a) - (b) saturation recovery images of water phantoms containing five different concentrations of $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}-1:2\text{w}:\text{Pluronic F-127}$ at two recovery delay times. Please note that the recovery at 10s is independent of the Fe ion content in each tube. The Fe ion content in each tube is shown in the inset; (c) T_1 relaxation map demonstrating the limited dependence of the longitudinal relaxation on the Fe ion content in phantoms; (d) - (e) spin echo images of water phantoms with the same Fe ion content as in (a) - (c); (f) corresponding T_2 map obtained using T_2 data analysis in each sample tube as described in the Experimental section. The following transverse

relaxation time constants are obtained after: (1) red - (0.239 ± 0.17) s; (2) yellow - (0.185 ± 0.010) s; (3) light green - (0.137 ± 0.010) s; (4) light blue - (0.108 ± 0.090) s; (5) dark blue - (0.072 ± 0.010) s; **(g)** transverse (red) and longitudinal (blue) relaxation rates as a function of iron content in the added contrast agent. Relaxivity is the slope extracted from linear fitting of the data; **(h)** comparison of r_2/r_1 ratio for commercial Combidex (C-dex), Feridex (F-dex), Resovist (R-vist)³⁶ and novel NP based contrast MRI agents³⁷ relative to this work at different magnetic field strengths.

The influence of the added contrast agent on the transverse proton relaxation is highlighted in Figure 6a. It is apparent that T_2 decreases with a corresponding increase of the iron concentration in the added contrast agent. On the contrary, the T_1 map displayed in Figure 6c shows mediocre dependence of the longitudinal relaxation on the iron content for the same range of concentration of the added contrast agent. To determine r_1 and r_2 relaxivities we analysed the dependence of the relaxation rates $r_1 (=1/T_1)$ and $r_2 (=1/T_2)$ on the iron ion concentration in the added contrast agent (**Figure 6g**). The relaxivities r_1 and r_2 were determined by linear fitting the data displayed in Figure 6g and resulted in $r_1 = (0.7 \pm 0.3)$ mM/s and $r_2 = (268 \pm 13)$ mM/s. The determined r_2/r_1 ratio is 380 ± 26 demonstrates that aqueous $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}-1:2\text{w}:\text{Pluronic F-127}$ surpasses the best commercially available agent (Resovist) by at least a factor of 6. This ratio is also higher (380 vs 306) than that reported for state-of-the-art magnetic nanoscale platforms³⁷ specifically designed for optimal performance at ultra-high (9.4T) magnetic field. This value highlights the great potential of $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}$ as “smart” contrast agents in *in vivo research* at ultra-high magnetic fields (7T and higher) as the majority of pre-clinical scanners operate at these field strengths.³⁸ With the further development of “smart” theranostic agents, like the one proposed here, the potential mechanism for personalised nanomedicine could be more efficiently explored in the pre-clinical setting.^{40,41} If successful, this could accelerate the pathway for bringing abdominal theranostics to the patient, of significant benefit due to the ongoing, worldwide rise of ultra-high field (7T) clinical MRI.^{42,43}

4. Conclusions

In conclusion, we have developed a simple one-step procedure to assemble $\text{Fe}_3\text{O}_4\text{NP}/\text{GNF}$ hybrid structures from widely available precursors. The influence of solvent, and the ratio of

starting complex to GNF, has also been studied. An optimised method using a 1:2 mass ratio of Fe(III) acetylacetonate to GNF, prepared using water as the solvent, showed excellent properties as an MRI contrast agent, outperforming commercially available materials and literature analogues. Whilst the issue of biocompatibility of carbon nanotubes and their composites with magnetite nanoparticles still needs further assessment,⁴⁴⁻⁴⁸ current literature studies suggest that analogous materials have limited toxicity to healthy and malignant cells and can be injected at a concentration of 2.5 mg of iron per kg of body weight without any apparent harm to health.^{49,50} Furthermore, the recent application of magnetic materials in CNTs for cell shepherding and magnetic fluid hyperthermia treatment⁵¹ represents a significant breakthrough for the safe use of these materials. Thus, when combined with the open internal channel of GNFs, accessible for therapeutic pharmaceutical agent loading alongside the Fe₃O₄ nanoparticles, such hybrid nanostructures can offer a novel mechanism for magnetically-induced drug release, in addition to the functional MRI contrast properties, consequently resulting in a potentially indispensable contribution to the emerging area of theranostics.

5. Acknowledgements

The authors are grateful to the Nanoscale & Microscale Research Centre (NMRC) for access to TEM and Raman facilities. This work supported by grant of Russian Federation Ministry of Science and Education Grant № K3-2015-030 (MISiS) and the University of Nottingham Advanced Molecular Materials for Healthcare Research Priority Area.

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