

Methane Adsorption in Metal-Organic Frameworks Containing Nanographene Linkers: a Computational Study

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Abstract

Metal-organic framework (MOF) materials are known to be amenable to expansion through elongation of the parent organic linker. For a family of model (3,24)-connected MOFs with the *rht* topology, in which the central part of organic linker comprises a hexabenzocoronene unit, the effect of the linker type and length on their structural and gas adsorption properties is studied computationally. The obtained results compare favourably with known MOF materials of similar structure and topology. We find that the presence of a flat nanographene-like central core increases the geometric surface area of the frameworks, sustains additional benzene rings, promotes linker elongation and the efficient occupation of the void space by guest molecules. This provides a viable linker modification method with potential for enhancement of uptake for methane and other gas molecules.

Keywords

metal-organic frameworks; grand canonical Monte Carlo simulations; methane storage; hexabenzocoronene

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I. Introduction

Crystalline nanoporous networks such as metal-organic frameworks (MOFs) with tuneable pore geometry and designed chemical functionality^{1,2} are attracting a great deal of interest as promising materials for a wide range of applications including light harvesting,^{3,4} drug delivery,⁵ biomedical imaging,⁶ catalysis,⁷⁻⁹ chemical sensing,¹⁰ gas separation and storage.¹¹⁻¹⁸ MOFs are compounds containing metal nodes (metal ions or clusters) coordinated to organic linker molecules that form the extended network structures with unique physical and chemical properties. The ability to tailor and control the density, internal pore volume and the internal surface area of MOFs^{19,20} is crucial for utilizing these materials in high capacity gas uptake applications. MOFs have shown great promise for mobile CH₄ sorption and storage,²¹ which has fuelled a further wave of interest in these materials, with the U.S. Department of Energy (DOE) launching a new CH₄ storage program²² with the ambitious new targets. In reference to the usable capacity stored between 35 and 5 bar at near-ambient temperature, the materials-level target for volumetric storage capacity of CH₄ after packing losses (25%) is 349 cm³/cm³ and 50 wt% for a gravimetric storage capacity. The drive to achieve these goals is powered by the current demand for alternative fuels and ever growing concerns over international energy security and climate change.

Methane storage in MOFs is developing rapidly,^{23,24} and several MOFs^{12,25-28} have been reported to show good volumetric capacity for CH₄ uptake at room temperature. It is highly desirable to increase an accessible surface area and introduce stronger interaction sites in order to obtain high CH₄ uptake. This can be achieved, for example, through chemical modification of organic linkers. A number of theoretical solutions have been offered previously for the enlargement of the surface area of MOFs and enhancement of gas uptake through modification of the linkers, notably by Duren *et al.*²⁹ for CH₄ storage and by Mavrandonakis *et al.*³⁰ for hydrogen storage. Wilmer *et al.*³¹ developed an exhaustive computational strategy for design and large-scale screening of hypothetical MOFs that allows the generation of new structures from a chemical “library of building blocks” based on the existing MOF structures. Over 300 MOFs with a predicted CH₄ storage capacity have been proposed using this structure generation approach. However, as geometry optimization procedure was not included, it is likely that some stable MOF structures with a high CH₄ uptake capacity were overlooked. Martin and Haranczyk have suggested³²⁻³⁴ an alternative strategy for computational design of the optimal organic ligand leading to an efficient occupation of the internal volume by guest molecules. In this approach, organic ligands are

replaced by purely geometrical “building blocks” described by a number of variable parameters, which were optimized to predict and iteratively refine the resulting MOF structure.

In this work, a new series of the (3,24)-connected $\{\text{Cu}_2(\text{COO})_4\}$ paddlewheel based MOF networks of *rht* topology has been designed in an effort to increase the surface area and pore volume. It has been found that the use of hexabenzocoronene unit as a central part of the linker leads to a significant enhancement uptake of CH_4 and other gases for the entire family of MOFs. This suggests that hexabenzocoronene is an excellent candidate for replacing the phenyl ring, a conventional central element of the linkers used in (3,24)-paddlewheel-connected MOFs.¹⁹ We have undertaken an exhaustive computational analysis to study the effect of the dimensions of the proposed hexabenzocoronene-based linkers on the gas storage capacity and other properties of this model family of MOFs, focussing especially on the surface area, pore volume, framework density and structure-property relationships. Particular attention was focused on the thermodynamic conditions at which the maximum gravimetric and volumetric uptake of CH_4 can occur. A similar computational strategy has been adopted previously by Fairen-Jimenez *et al.*³⁵ in the study of hydrogen adsorption in hypothetical MOFs with the *rht*-topology.

II. Structural properties

The structural models are based on the (3,24)-connected network with *rht* topology previously used by Yan *et al.*,^{36,37} Nouar *et al.*,³⁸ Yuan *et al.*,^{39,40} and Farha *et al.*^{19,20} *Rht*-topology has been selected to avoid the interpenetration and/or interweaving that greatly affects the gas sorption properties. The new structures contain vacant areas at each of the Cu(II) sites within the binuclear paddlewheel nodes, which promote binding between the metal and the adsorbate gas molecules. We use the asymmetric unit of **NOTT-112**³⁷ (Figure 1a) to form the Cu(II) paddlewheel cluster within the *rht* network topology with *Fm-3m* group symmetry, and replace the central phenyl core with hexabenzocoronene moiety (Figure 1b) to promote the linker elongation and efficient occupation of the void space by guest molecules. As shown in Figures 1a,b the replacement of the central part in the linker of **NOTT-112** with hexabenzocoronene leads to an increase of its dimensions by almost 60%. Seven different linker fragments, labelled **L¹-L⁷** (Figure 1c), have been used with the hexabenzocoronene central element to construct a family of frameworks with *rht*-network topology (see Figure S1 in *Supporting Information*).

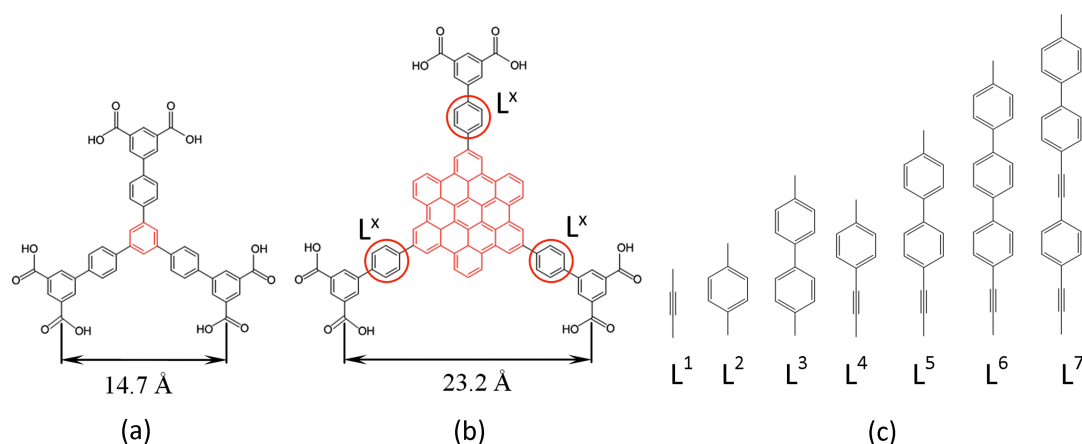


Figure 1: Views of (a) organic linker of **NOTT-112**,³⁷ (b) organic linker containing hexabenzocoronene and the L^x fragment; (c) the L^1 - L^7 linker fragments.

An additional framework, labelled **MOF- L^0** , has been built from a hexacarboxylate linker where three isophthalate units are directly connected to the central hexabenzocoronene part. The structures of the designed frameworks have been optimized using molecular mechanics, and their dimensions have been compared to un-modified MOFs containing a phenyl ring as the central core of the linker (see *Supporting Information* for details). The structure of a member of the model MOF family, **MOF- L^1** , containing hexabenzocoronene and the shortest organic linker L^1 , is shown in Figure 2.

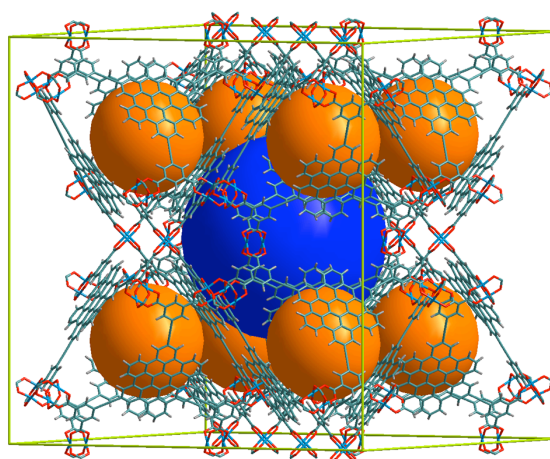


Figure 2: View of the structure of model **MOF- L^1** containing hexabenzocoronene and the shortest organic linker L^1 .

Replacement of the central part of the hexacarboxylate with hexabenzocoronene in the (3,24)-connected MOF networks leads to a significant change not only in the dimensions but also in other structural properties of the proposed family of model MOFs. A summary of the

structural properties of the MOFs family containing hexabenzocoronene central part, which include the geometric surface area, pore volume, framework density and unit cell length, is given in Table 1.

Table 1: The structural properties of the model MOFs family containing hexabenzocoronene molecule as a central linker fragment.

	geometric surface area, * (m ² /g)	pore volume, (cm ³ /g)	framework density, (g/cm ³)	unit cell length, (Å)
MOF-L⁰	2589	1.33	0.61	47.01
MOF-L¹	4172	1.88	0.46	52.93
MOF-L²	5048	2.40	0.36	59.44
MOF-L³	5784	3.19	0.28	65.98
MOF-L⁴	5784	3.50	0.26	70.06
MOF-L⁵	6497	4.52	0.20	76.93
MOF-L⁶	7014	6.09	0.15	87.93
MOF-L⁷	7514	7.42	0.12	94.68

*geometric surface area is calculated using Materials Studio 4.4

The Brunauer-Emmett-Teller (BET) surface area of **NOTT-112** is reported to be 3800 m²/g,³⁷ whereas the geometric surface area of **MOF-L²** is 5048 m²/g, approximately 33% increase upon modification of the central element of the linker. The model **MOF-L³** framework, which has been obtained by replacing the central element of the **NOTT-119**³⁶/**PCN-69**⁴⁰ linker with hexabenzocoronene, also exhibits an enhanced geometric surface area of 5784 m²/g, which compares favourably with the surface area of un-modified **NOTT-119** (BET surface area of 4118 m²/g)³⁶ and **PCN-69** (BET surface area of 3989 m²/g).⁴⁰ Similar increase in the surface area has been achieved in **MOF-L⁴** (5784 m²/g), which structure can be compared to that of the un-modified **NOTT-116** (BET surface area of 4664 m²/g)³⁶ and **PCN-68** (BET surface area of 5109 m²/g)³⁹ MOFs, and in **MOF-L¹** (4172 m²/g), which can be compared to the un-modified **PCN-61** (3500 m²/g)³⁹ MOF.

An alternative way to compare the proposed model family of MOFs with existing un-modified MOFs structures is shown in Figures 3a-c. The un-modified structures (right panel of Figure 3) are presented against the modified model MOFs with the identical central part (highlighted in blue, left panel), which was decorated by additional nine benzene rings. The optimised length of the cubic unit cell of **MOF-L²** is calculated to be 59.44 Å, which is comparable to the unit cells of the MOF structure designed by Fairen-Jimenez *et al.* using the **L³** fragment³⁵ (58.32 Å), **NOTT-119**³⁶ (56.30 Å) and **PCN-69**⁴⁰ (59.15 Å simulation; 56.61 Å

experiment). A discrepancy between the model and experimental values of the length of unit cell might be due to the fact that organic linkers in the computationally designed MOFs remain predominantly straight, while the linkers of **PCN-69**⁴⁰ and **NOTT-119**³⁶ bend to a certain extent.⁴⁰ The geometric surface areas of the model MOFs are also similar: 5049 m²/g in **MOF-L²** and 5194 m²/g in the model MOF designed by Fairen-Jimenez *et al.* using the **L³** fragment,³⁵ however the BET surface area in the synthesized MOFs is somewhat smaller, *e.g.* 4118 m²/g in **NOTT-119**,³⁶ and 3989 m²/g in **PCN-69**.⁴⁰

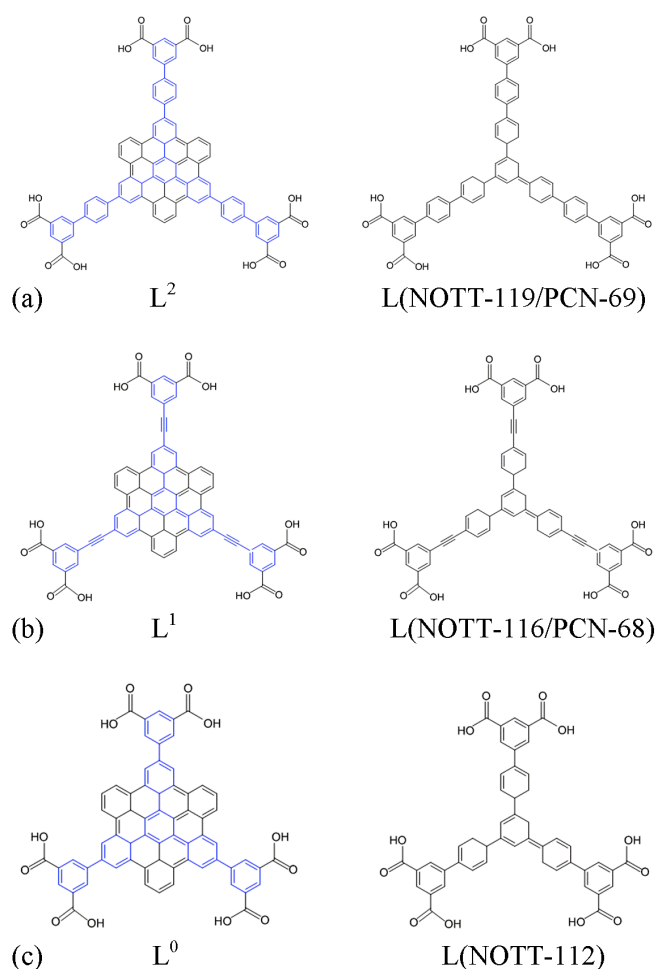


Figure 3: Comparison of linker geometries of three members of the proposed model MOFs family, based on hexabenzocoronene molecule, with existing experimental MOFs: a) model **MOF-L²** and **NOTT-119**,³⁶ b) model **MOF-L¹** and **NOTT-116**,³⁶ c) model **MOF-L⁰** and **NOTT-112**.³⁷

The model **MOF-L¹** framework shown in Figure 3 has a basic structure of the existing unmodified frameworks **NOTT-116**³⁶ and **PCN-68**³⁹ and the model MOF designed by Fairen-Jimenez *et al.* using **NOTT-112** and the **L⁴** fragment.³⁵ The optimised length of the unit cell of **MOF-L¹** is predicted to be 52.93 Å, in good agreement with the available experimental (52.74 Å in **PCN-68**)³⁹ and computational data (53.73 Å in the model MOF³⁵ based on

NOTT-112 and **L⁴** fragment). Such good agreement between the computational and experimental data is observed because the organic linkers in these MOFs are entirely straight. However, the predicted geometric surface area in **MOF-L¹** (4172 m²/g) is smaller than the BET surface area in **NOTT-116** (4664 m²/g),³⁶ BET surface area in **PCN-68** (5109 m²/g)³⁹ and the geometric surface area in the model MOF designed by Fairen-Jimenez *et al.* using **NOTT-112** and the **L⁴** fragment (5033 m²/g).³⁵ Finally, the smallest **MOF-L⁰** framework can be directly compared to **NOTT-112**. The unit cell length of 47.01 Å in **MOF-L⁰** is exactly the same as the one experimentally obtained for **NOTT-112**,³⁷ although the geometric surface area of **MOF-L⁰** (2589 m²/g) is lower than the BET surface area of **NOTT-112** (3800 m²/g).³⁷ The comparison method used in Figure 3 shows MOF structures of very similar dimensions, but the estimation of their accessible surface area depends very strongly on the method and type of a probe used.

The largest MOF of the proposed family, **MOF-L⁷**, exhibits a surface area of 7514 m²/g, which exceeds the largest BET surface area reported in experiment to date in this series, that of **NU-109** (7000 m²/g).¹⁹ The **MOF-L⁷** framework density of 0.12 g/cm³ is also at the limit close to the lowest calculated density for porous crystals reported for the **MOF-399** compound (0.126 g/cm³).⁴¹ The remaining members of the proposed family exhibit structural properties comparable with typical values achieved in the synthesis of ultrahigh porosity MOFs.⁴² A longer organic linker typically provides larger void space and a greater number of adsorption sites within a framework. However, linkers with a very large number of phenyl repeat units make a MOF liable to structure interpenetration, poor solubility, low synthetic yields, and cumbersome purification protocols. The use of MOFs with the *rht* topology, based on a singular net for the combination of 3- and 24-connected nodes, removes any concern for catenation (interpenetration or interweaving of multiple frameworks) and makes them an ideal target in design of ultrahigh porosity. The desolvation and associated activation using supercritical CO₂ and solvent exchange and processing^{19,43} is a particularly promising method that allows generation of MOFs with ultrahigh surface areas.

III. Gas uptake performance

The composition, structure and density of the organic linker, not only its length, also play an important role in design of high surface area and high gas uptake MOFs. Farha *et al.*⁴⁴ showed that replacing the phenyl rings of organic linkers with triple-bond spacers is another effective method of improving gravimetric surface areas and, hence, porosity of MOFs. A

comparison of the structural properties of the L^3 and L^4 linker fragments, which contain two phenyl spacers and a phenyl spacer and a triple bond, respectively, supports this observation. Indeed, both $\text{MOF-}L^3$ and $\text{MOF-}L^4$ have the same geometric surface area of $5784 \text{ m}^2/\text{g}$ (Table 1) and a very close gas uptake performance at a range of temperatures (see Figure 5 below). However, phenyl rings are found to be quite strong adsorption sites. The “resolution of the identity” integral approximation applied to second-order many-body perturbation theory (RI-MP2), as implemented in the Q-Chem quantum chemistry package,⁴⁵ has been employed to analyse the strength of the binding between a phenyl ring and a methane molecule. The RI-MP2 method offers improved computational performance compared to traditional exact second-order perturbation theory (MP2) calculations. A phenyl ring of the linker structure is represented as a benzene molecule; and the geometry optimizations of the benzene – methane dimer and corresponding monomers have been obtained at the RI-MP2/cc-pvqz level of theory. The binding energies were corrected using basis set superposition error (BSSE).⁴⁶ The most stable interaction geometry is found to be the ANTI $\text{C}_6\text{H}_6 - \text{CH}_4$ dimer, which has the binding energy of -7.30 kJ/mol (with BSSE correction), and it is shown in Figure 4. The obtained values for the binding energy of the ANTI $\text{C}_6\text{H}_6 - \text{CH}_4$ dimer is in a good agreement with the RI-MP2/QZVPP value of -7.01 kJ/mol .⁴⁷ Although the calculated binding energy gained from forming ANTI $\text{C}_6\text{H}_6 - \text{CH}_4$ dimer is weaker than a typical weak hydrogen bond like interaction it is sufficient to surpass the kT energy barrier (at room temperature).

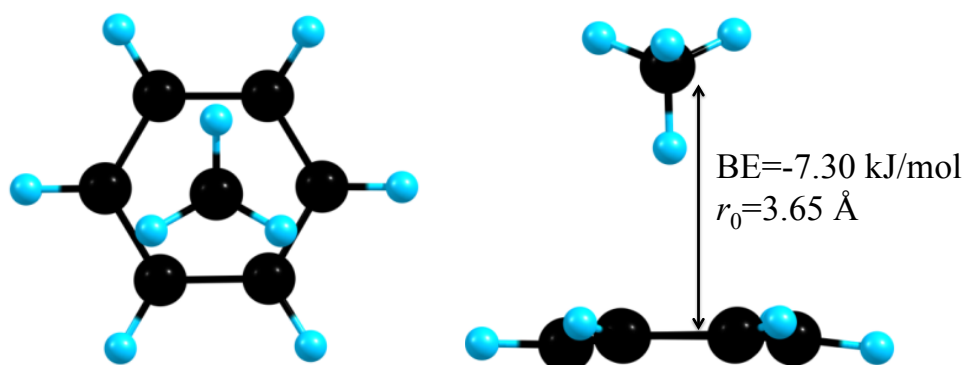


Figure 4: The most stable interaction geometry, ANTI $\text{C}_6\text{H}_6 - \text{CH}_4$, as predicted by the RI-MP2/cc-pvqz: the BSSE corrected binding energy is $\text{BE} = -7.30 \text{ kJ/mol}$ and the equilibrium bond distance, defined as the distance between the centre of mass of each molecule, is $r_0 = 3.65 \text{ \AA}$.

Grand Canonical Monte Carlo (GCMC) simulations have been used to predict CH_4 uptake in the model MOF structures at $T = 273 \text{ K}$ and $T = 298 \text{ K}$ and for pressures up to 70 bar (details of the simulation approach can be found in *Supporting Information*). The MOF

frameworks and the guest gas molecules were considered to be rigid. The Lennard-Jones (LJ) potential was used to describe the Van der Waals interactions with a cut-off distance of 12.8 Å. The values of the LJ parameters for atoms present in MOFs are summarised in Table S2 of *Supporting Information*. Most of the LJ parameters were taken from the DREIDING force field, which uses general force constants and geometric parameters based on simple hybridisation considerations, so that the bond distances are derived from atomic radii, and there is one force constant each for bonds, angles, and inversions and six values for torsional barriers. The LJ parameters for copper atoms were taken from the Universal Force Field. Both the total and excess gravimetric and volumetric sorption isotherms of the model MOFs for CH₄ uptake were calculated, and the total uptake data are presented in Figure 5 (the excess adsorption isotherms for CH₄ in model MOFs with the L⁰-L⁷ organic linkers at T = 273, 298 K are shown in Figure S5 of *Supporting Information*). The summary of sorption data at 35 and 70 bar is given in Table 2. At 273K and 70 bar, the model MOF structures with the L³-L⁷ linker fragments show a significant total gravimetric uptake exceeding 50 wt% (wt% is equal to amount of adsorbed CH₄ relative to mass of framework). The model **MOF-L⁷** with the longest linker exhibits the highest absolute gravimetric uptake of 78 wt%. Although **MOF-L⁷** has extremely large pore openings, the dimensions of **MOF-L³** are comparable to those of **NOTT-119**³⁶/**PCN-69**⁴⁰, which have been synthesized successfully (see details in *Supporting Information*). As the computational isotherms correspond to a perfect structure with no impurities or pore collapse, the uptake values for CH₄ shown in Figure 5 can only be considered as the upper limit. Various structural defects or collapsed regions that may occur naturally in experimental samples will lower these theoretical estimates. On a volumetric basis at high pressures the structures exhibit the opposite trend so that the lowest volumetric uptake of CH₄ (138.24 cm³/cm³ at T = 273 K and 70 bar) corresponds to the structure with the longest linker, **MOF-L⁷**.

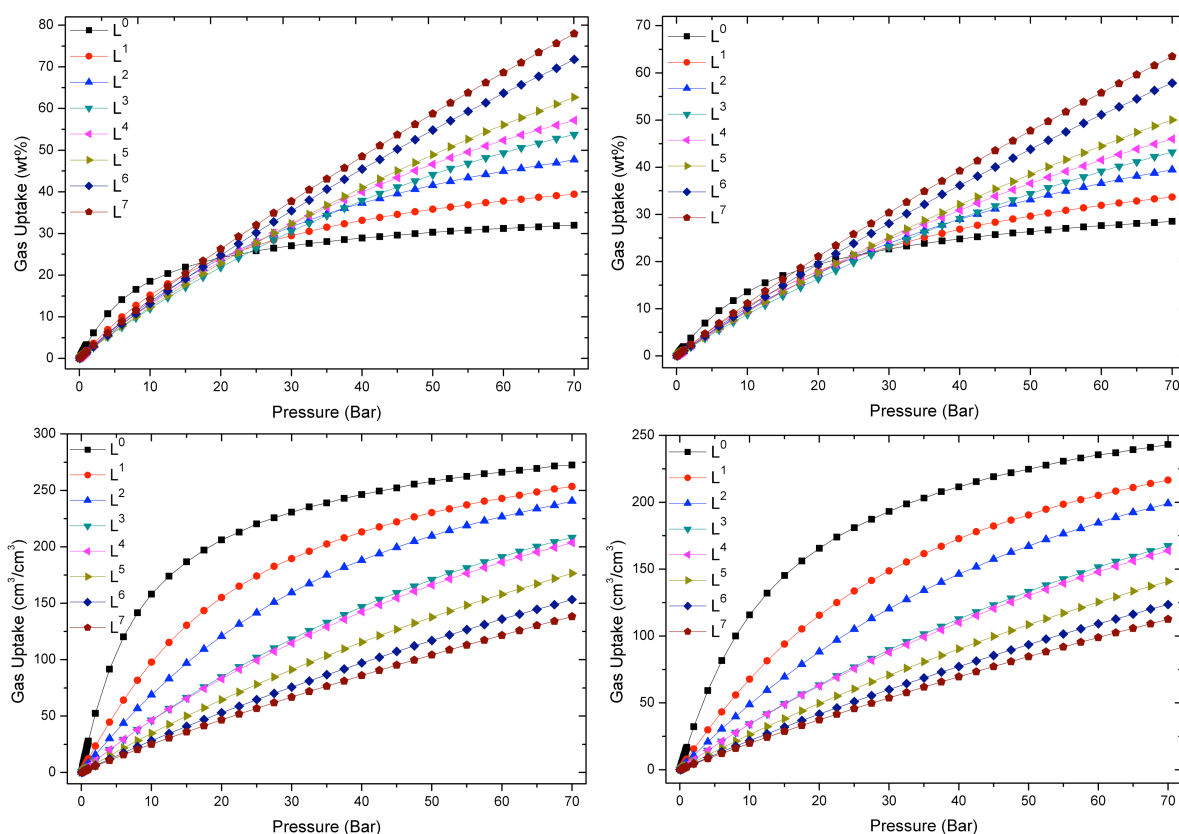


Figure 5: Calculated total adsorption isotherms for CH₄ in model MOFs with the L⁰-L⁷ organic linkers: gravimetric (top) and volumetric (bottom) results at T=273 K (left) and T=298 K (right).

On the other hand, the framework with no linker fragment, L⁰, exhibits an impressive total volumetric uptake of CH₄ of 273 cm³/cm³ at 273 K and 70 bar, although the total gravimetric uptake of this MOF (32 wt%) is the lowest in this series under the same conditions. At 298 K, the overall trend remains the same confirming that compared to the MOFs with longer linkers, the model MOFs with shorter linkers achieve higher volumetric but lower gravimetric sorption at high pressures. Due to the nature of the physical adsorption of gases in MOFs, the gravimetric uptake depends directly on the surface area and pore volume. Thus, porous materials with high pore volumes show high gravimetric gas uptake at saturation; the larger the gas cylinder the more gas can be stored at saturation. However, structures with high surface areas often possess low framework densities, and thus, in volumetric terms (cm³ of gas per cm³ of host), such highly porous materials often show lower volumetric gas capacities,⁴⁸ as observed in the current study. Therefore, a balance of surface area, pore volume and framework density are important factors in designing materials for optimum volumetric or gravimetric CH₄ adsorption.⁴⁹

MOFs with the linkers L¹ and L² underperform systematically in both volumetric and gravimetric metrics. An increase in temperature at constant pressure reduces the overall CH₄

uptake performance so that only three structures exceed the total gravimetric uptake of 50 wt%, namely **MOF-L⁷** with 63.5 wt%, **MOF-L⁶** with 59 wt% and **MOF-L⁵** with 50 wt% total gravimetric uptakes (see Table 2 for further details). The volumetric uptake also decreases with increase of temperature in all considered model MOF structures. It can be generally observed at 298 K and 70 bar that the model MOFs **L²-L⁷** show higher gravimetric uptake than **NU-111** (36 wt%),⁴⁹ **PCN-68** (35 wt%)³⁹ and **NOTT-119** (30 wt%)³⁶ reported in the literature. The observed strong temperature dependence on CH₄ uptake may potentially provide an efficient tool for gas uptake and subsequent release.

Table 2: Summary of CH₄ sorption characteristics in **MOF-L⁰** - **MOF-L⁷** calculated at T = 273 K and T = 298 K and the pressures of 35 bar and 70 bar; v/v denotes cm³/cm³.

	70 bar								35 bar								working capacity at 298 K*	
	T=273 K				T=298 K				T=273 K				T=298 K					
	total uptake		excess uptake		total uptake		excess uptake		total uptake		excess uptake		total uptake		excess uptake		wt%	v/v
	wt%	v/v	wt%	v/v	wt%	v/v	wt%	v/v	wt%	v/v	wt%	v/v	wt%	v/v	wt%	v/v		
MOF-L⁰	32.0	273	24.2	206	28.5	243	21.7	185	28.0	239	24.4	208	23.8	203	20.6	176	20.4	174
MOF-L¹	39.4	254	28.3	182	33.7	216	24.1	155	31.5	203	26.4	170	25.1	162	20.6	132	28.1	181
MOF-L²	44.7	241	30.6	154	39.5	199	27.3	137	34.7	175	28.2	142	26.6	134	20.8	105	34.1	172
MOF-L³	53.7	208	34.9	137	43.2	167	26.9	105	34.3	133	25.7	101	26.2	102	18.6	73	38.6	151
MOF-L⁴	57.2	204	36.6	133	46.0	164	28.1	102	36.3	129	26.9	98	27.9	99	19.5	71	40.9	149
MOF-L⁵	62.7	177	36.1	101	50.0	141	26.9	75	36.8	104	24.6	69	28.7	81	17.8	50	45.2	126
MOF-L⁶	71.8	153	36.0	75	57.9	124	26.8	56	40.6	87	24.2	51	32.2	69	17.5	37	52.4	110
MOF-L⁷	78.0	138	34.3	58	63.5	113	25.6	43	43.1	76	23.1	39	34.9	62	17.0	29	57.6	97

*the working capacity is defined as the difference in total uptake between 70 bar and 5 bar.

Table 2 and Figure 5 confirm that for a constant temperature CH₄ uptake is lower at 35 bar than at 70 bar, as expected. At 35 bar and 273 K the model **MOF-L⁷** framework shows the highest gravimetric uptake of 43.1 wt%, whereas the highest volumetric uptake of 239 cm³/cm³ is shown by **MOF-L⁰**.

Table 3 presents a direct comparison of the gravimetric and volumetric excess CH₄ uptake performance at 298K and 35 bar in the model **MOF-L¹** and **MOF-L³** frameworks

modified with the hexabenzocoronene central unit and the corresponding un-modified MOFs reported in the literature, **PCN-61**³⁹ and **PCN-68**³⁹. The modified model **MOF-L**¹ shows an excess uptake of methane of 132.4 cm³/cm³ and 20.6 wt%, which is significantly higher than the excess uptake values of 99 cm³/cm³ and 18.6 wt% in the un-modified **PCN-68** MOF³⁹ with the same dimensions as **MOF-L**¹ (Figures 3b and S2 of *Supporting Information*). However, a larger **MOF-L**³ framework exhibits an excess methane uptake of 72.8 cm³/cm³ and 18.6 wt%, which is only comparable to the methane uptake performance in smaller **PCN-68** framework. The predicted values for the total and excess gravimetric and volumetric CH₄ uptake performance at 298K of the larger model **MOF-L**⁵- **MOF-L**⁷ frameworks are also comparable to those reported in the ultrahigh surface area **NU-100/NU-109** MOFs.⁵⁰

Table 3: Comparison of the gravimetric and volumetric excess CH₄ uptake performance at 298K and 35 bar in the model **MOF-L**¹ and **MOF-L**³ frameworks modified with the hexabenzocoronene central unit and the corresponding un-modified MOFs reported in the literature, **PCN-61**³⁹ and **PCN-68**³⁹.

	MOF-L ¹	MOF-L ³	PCN-61 ³⁹	PCN-68 ³⁹
gravimetric excess uptake, wt%	20.6	18.6	18.6	18.6
volumetric excess uptake, cm ³ /cm ³	132.4	72.8	145.0	99.0

When considering the performance of a material for practical use in automobile application, the working capacity, which defines how far a car can travel, is more relevant than the adsorption capacity. The working capacity, defined as the deliverable amount of methane from two different pressures (70 bar to 5 bar), is shown in Table 2. The gravimetric working capacity increases with the size of the linker in this series of frameworks. This is consistent with the fact that the MOFs with higher surface areas and pore volumes can hold more CH₄ at high pressures. **MOF-L**¹ with an optimized pore size and surface area shows the highest volumetric deliverable capacity of 180.6 cm³/cm³ at 298 K.

GCMC simulations have also been performed to predict CO₂ and H₂ sorption in the largest framework of the considered model MOF family, **MOF-L**⁷.

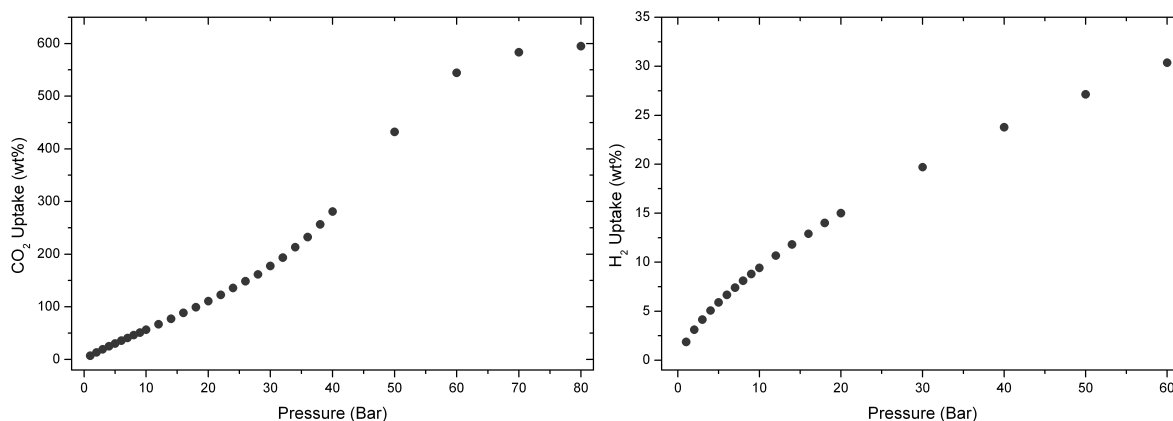


Figure 6: Adsorption isotherms in the model MOF with the L^7 organic linker calculated for CO_2 (left) at $T = 298$ K and H_2 (right) at $T = 77$ K.

Simulation of CO_2 sorption (Figure 6) reveals a high CO_2 storage capacity for **MOF- L^7** at 298 K, 80 bar of 595 wt%. Due to its very large pore size **MOF- L^7** is not suitable for the CO_2 sorption at low pressures, but at high pressures a stepwise behaviour in the adsorption isotherm is observed, and a significant amount of CO_2 is adsorbed in the structure. Therefore, pressures suitable for CO_2 storage in these structures must be higher than 70 bar. The highest experimental values for total CO_2 uptake achieved so far have been reported for **NU-100** (232 wt%)²⁰ at 40 bar and 298 K and **MOF-210** (248 wt%)⁴⁹⁻⁵¹ at 50 bar and 298 K. The total uptake of 30.4 wt% for H_2 gas in **MOF- L^7** has been obtained at 60 bar and 77 K, and is a twice as high as the than the known experimental values for the best performing MOFs with **NU-100** and **MOF-210** showing total H_2 uptakes of 16.4 wt%²⁰ and 16.7 wt%,^{49,51} respectively.

V. Conclusions

A family of model MOFs based on the (3,24)-connected MOF networks with the *rht* topology has been proposed in which the central part of the organic linker has been replaced with hexabenzocoronene molecule. An exhaustive computational analysis allows not only prediction of the physical and chemical properties of the proposed MOFs family, but also a direct comparison of the model structures with existing MOFs of similar structure. It has been shown that a replacement of the central linker fragment with hexabenzocoronene molecule increases the size and, in some cases, gas uptake capability of the MOFs family with *rht*-network topology. The presence of a small flat nanographene-like central part increases the geometric surface area not only through the provision of additional benzene rings, but also by providing some extra rigidity to the linker structure.

GCMC simulations of CH₄ uptake have been performed on the optimised MOF structures at pressures up to 70 bar and at two temperature values, T = 273 K and 298 K. This study confirms that a number of the designed MOFs formally reach the DOE targets²² for CH₄ storage, but none of the proposed MOFs achieves both gravimetric and volumetric uptake targets. In addition to CH₄ storage capability, the GCMC simulations of CO₂ and H₂ have been performed for the largest family member, **MOF-L**⁷, which exhibits outstanding sorption properties at high pressures. Overall, it has been demonstrated that a replacement of the central fragment with hexabenzocoronene molecule is a viable linker modification method, which might have the potential in enhancement of gas uptake for CH₄ and other gaseous molecules. derived from hexabenzocoronene has not been reported thus far, although precursors to such materials are known, and significantly there is a very recent report of porous organic materials derived from triiodo-hexabenzocoronene.⁵² This suggests that MOFs derived from and incorporating the hexabenzocoronene core are realistic synthetic targets.

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Supporting Information Available

Additional structural information on the model family of the (3,24)-paddlewheel-connected MOF networks with *rht* topology containing hexabenzocoronene molecule and further computational details are included in Supporting Information. This information is available free of charge via the Internet at <http://pubs.acs.org>.

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Graphic with Article submission

